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Cool! I'am really happy

#Markus Jensen



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#Diego Butler



so many fake sites. this is the first one which worked! Many thanks

A concentrated solution of aqueous aluminum sulfate is made by dissolving 274 g solid aluminum sulfate in enough water to make 975 mL of solution. Then 7.50 mL of this concentrated solution is added to 60.0 mL of water. Calculate the sulfate ion concentration in the final dilute solution.

Step 1: Calculate sulfate ion concentration of concentrated solution

$$\text{Molarity} = \frac{\text{moles solute}}{\text{L solution}} = \frac{2.40 \text{ mol SO}_4^{2-}}{0.975 \text{ L solution}} = 2.46 \text{ M SO}_4^{2-}$$

Find moles SO_4^{2-} :

$$274 \text{ g Al}_2(\text{SO}_4)_3 \times \frac{1 \text{ mol Al}_2(\text{SO}_4)_3}{342.17 \text{ g Al}_2(\text{SO}_4)_3} \times \frac{3 \text{ mol SO}_4^{2-}}{1 \text{ mol Al}_2(\text{SO}_4)_3} = 2.40 \text{ mol SO}_4^{2-}$$

Find L solution:

1 L solution

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How To Calculate Ion Concentration In Solution