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Question

During the purification of iron in a blast furnace, the impurity tetraphosphorus decoxide,  $P_2O_5(l)$ , reacts with solid calcium oxide to produce molten calcium phosphate, or slag.

Write the balanced equation for this reaction. States are graded.

If the furnace initially contained 20.7 kg of solid calcium oxide and 11.5 kg  $P_2O_5(l)$ , how many kilograms of calcium phosphate were produced during the purification of the iron?

Number  kg

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